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Synthesis And Bio-Spectral Studies of Ni(II) Complex of 2'-Hydroxy-4'-Methoxyacetophenoneoxime (HMAOX)

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ABSTRACT

Ni(II) complex of 2'-hydroxy-4'-methoxyacetophenoneoxime (HMAOX) was synthesized from paeonol oxime by using standard protocol, and characterized by elemental analyses, melting point determination and spectral data. The ML₂ (metal/ligand) stoichiometry of the complex was determined by spectrophotometric and potentiometric studies, and mass spectral data. The value of stability constant of the complex was found to be 2.36×10^8 while its standard free energy of formation is 11.436 kcal mol⁻¹ at 27°C. Beer's law is obeyed in the concentration range 25-140 ppm of Ni. The value of molar extinction coefficient and sensitivity as per Sandell's scale were found to be 4.56×10^3 L.mol⁻¹ cm⁻¹ and 0.128 µg Ni/cm² respectively. The IR studies reveal that the phenolic proton is lost on complexation and the oxygen of the phenolic (-OH) and nitrogen of the oximino (=NOH) groups coordinate with Ni(II) ion. The electronic spectra and magnetic susceptibility measurement indicate that the complex is diamagnetic and square-planar in nature. The antimicrobial activity of different concentrations of ligand and its Ni(II)-complex has been evaluated against Aspergillus niger, Aspergillus flavus, Aspergillus nidulans and Alternaria alternata fungi and Staphylococcus, Streproproteus, Staph and Escherchia coli bacteria. The results indicated that the ligand (HMAOX) and its Ni(II) complex have good anti-microbial properties as compared to the standard drugs (fluconazole and ciprofloxacin). The activity index (AI) for the, bioactivity was also derived.

Keywords: Ni(II)-complex, Spectra, Thermodynamic parameters, Antimicrobial screening.