



Kinetics of Ir(III) Catalyzed Oxidative Decarboxylation and Deamination of L-phenylalanine by Chloramine-T: A Mechanistic Approach

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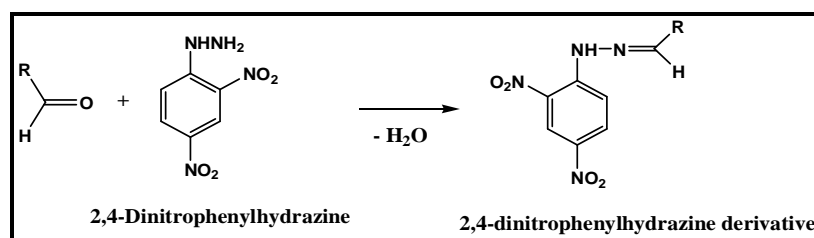
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ABSTRACT

The kinetics and mechanism of homogeneously Ir(III) chloride catalyzed oxidation of L-Phenylalanine by chloramine-T [CAT] has been investigated in perchloric acidic medium in presence of mercuric acetate as a scavenger in the temperature range 30°C-45°C. The experimental results show first order kinetics with respect to the oxidant [CAT] and catalyst [Ir(III)] while positive effect with respect to substrate i.e. L-Phenylalanine was observed. The reaction shows negligible effect of [Hg(OAc)₂], [H⁺] and ionic strength(μ)of the medium. Chloride ion positively influenced the rate of reaction. The reaction between chloramine-T and substrate (L-Phenylalanine) in acid medium shows 1:1 stoichiometry. To calculate activation parameters, the reactions have been studied at four different temperatures between 30 to 45°C. A mechanism involving the complex formation between catalyst, substrate and oxidant has been proposed. Phenyl acetaldehyde has been identified chromatographically and spectroscopically as the final product of oxidation of L-Phenylalanine. Based on the kinetic data, reaction stoichiometry and product analysis, a reaction mechanism has been proposed and rate law has been derived.

Graphical Abstract



Preparation of 2, 4-dinitrophenyl hydrazine derivative

Keywords: Kinetics, Mechanism, Ir(III) catalysis, Amino acids, Chloramine-T, Acidic medium.